

Transformational Environmental

Technologies

UNSW Centre for UNSW Centre for Transformational Environmental Technologies

# **Catalytic Ozonation Reactor Design** using Computational Fluid Dynamics (CFD) modelling

## **Our expertise:**

Using the state of art technology, CFD coupling with chemical reactions to optimise the catalytic ozonation reactor design.

World-class Modelling

**High Efficiency** 

**High Precision** 

### **Catalytic Ozonation in Full-scale Reactors**

A set of **elementary chemical reactions** with constant kinetic rates. Reactions in full-scale reactors are complicated, involving the interactions between fluid flow, mass transfer and chemical reactions. Similar to the three-dimensional distribution of flow field in the full-scale reactors, the reactants and products are usually **unevenly distributed** in the reactors, leading to the consideration of 'Chemical Reaction Dead Zone' in addition to 'Flow Dead Zone' in catalytic ozonation reactor design.



## 1 Kinetic Modelling of Chemical reactions

Traditional experimental methods was employed to obtain the kinetic models for pure ozonation of formate.

#	Elementary reaction	Reaction rate
1	$0_3 \rightarrow H_2 O_2 + O_2$	k <sub>1</sub> =2.21e-4 [s <sup>-1</sup> ], R = k <sub>1</sub> [O <sub>3</sub> ]
2	$O_3 + H_2 O_2 \rightarrow OH^* + O_2 + O_2^*$	k <sub>2</sub> =97.6 [L/mol/s], R = k <sub>2</sub> [O <sub>3</sub> ] [H <sub>2</sub> O <sub>2</sub> ]
3	$0_3 + 0_2^* \rightarrow 0H^* + 0_2$	k <sub>3</sub> =1.5e+09 [L/mol/s], R = k <sub>3</sub> [O <sub>3</sub> ] [O <sub>2</sub> *]
4	$\mathcal{CO}_3^* + \mathcal{O}_3 \to \mathcal{HCO}_3^- + \mathcal{O}_2$	k <sub>4</sub> =1.0e+05 [L/mol/s], R = k <sub>4</sub> [CO <sub>3</sub> <sup>2-</sup> ] [O <sub>3</sub> ]
5	$OH^* + HCO_3^- \rightarrow CO_3^* + OH^-$	k <sub>5</sub> =1.22e+07 [L/mol/s], R = k <sub>4</sub> [OH*] [HCO <sub>3</sub> <sup>-</sup> ]
6	$CO_3^* + H_2O_2 \rightarrow HCO_3^- + O_2^*$	$k_6 = 4.30e + 05 [L/mol/s], R = k_6 [H_2O_2] [CO_3^{2-}]$
7	$HCOOH + O_3 \rightarrow CO_3^* + O_2^*$	k <sub>7</sub> =10[L/mol/s], R = k <sub>4</sub> [HCOOH] [O <sub>3</sub> ]
8	$HCOOH + CO_3^* \rightarrow CO_2^* + HCO_3^-$	k <sub>8</sub> =1.5e+05[L/mol/s], R = k <sub>8</sub> [HCOOH] [CO <sub>3</sub> *]
9	$HCOOH + OH^* \rightarrow CO_2^* + H_2O$	k <sub>9</sub> =3.2e+09[L/mol/s], R = k <sub>9</sub> [HCOOH] [OH*]
1 0	$CO_2^* + O_2 \rightarrow HCO_3^- + O_2^*$	k <sub>10</sub> =4.2e+09[L/mol/s], R = k <sub>10</sub> [CO <sub>2</sub> *] [O <sub>2</sub> ]

#### 2 Coupling of Hydrodynamics and reactions



Gas-phase ozone mass fraction



Dissolved ozone concentration





#### **3** Continuous Flow Experiments for Model validation





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